



The Explosive Lives of Stars: Producing Elements in the Cauldrons of the Cosmos

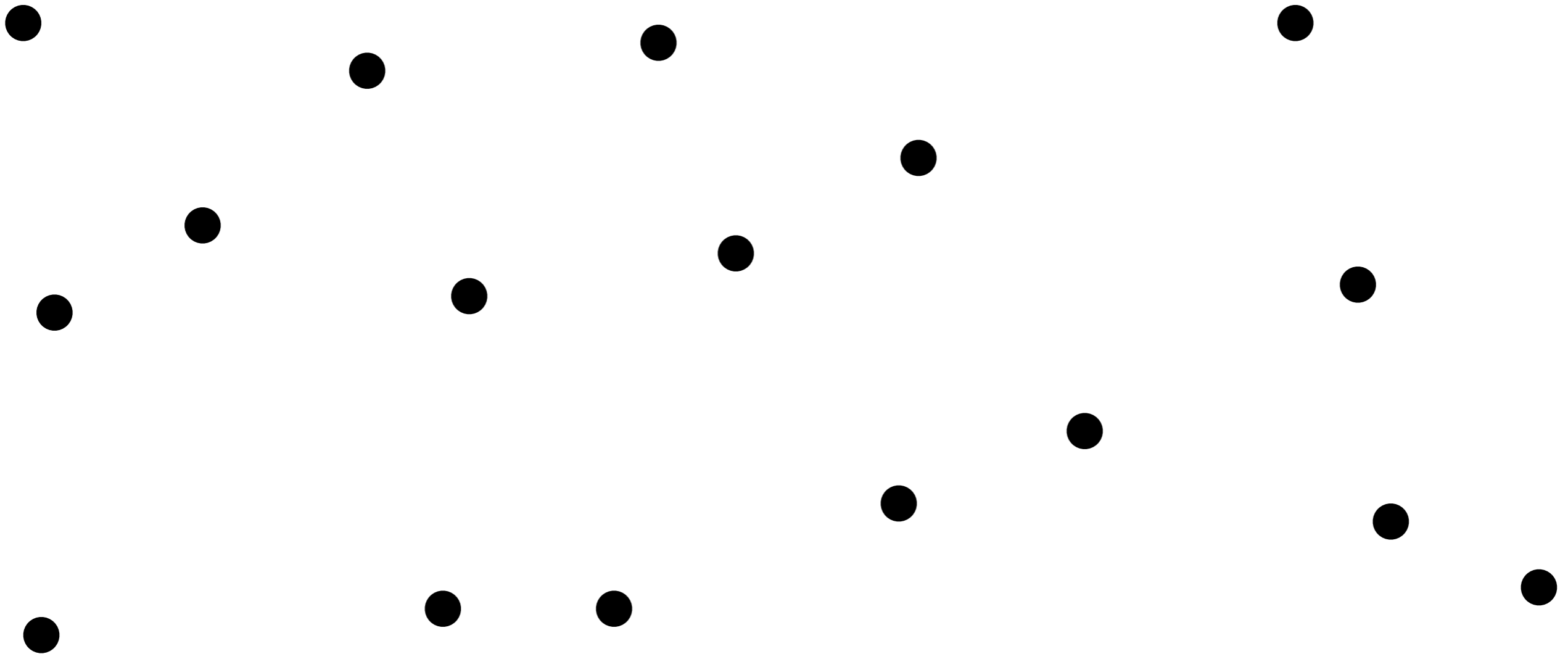
Catherine M. Deibel
Department of Physics & Astronomy
Louisiana State University

13.7 Billion (13,700,000,000) years ago the
universe began with

THE BIG BANG

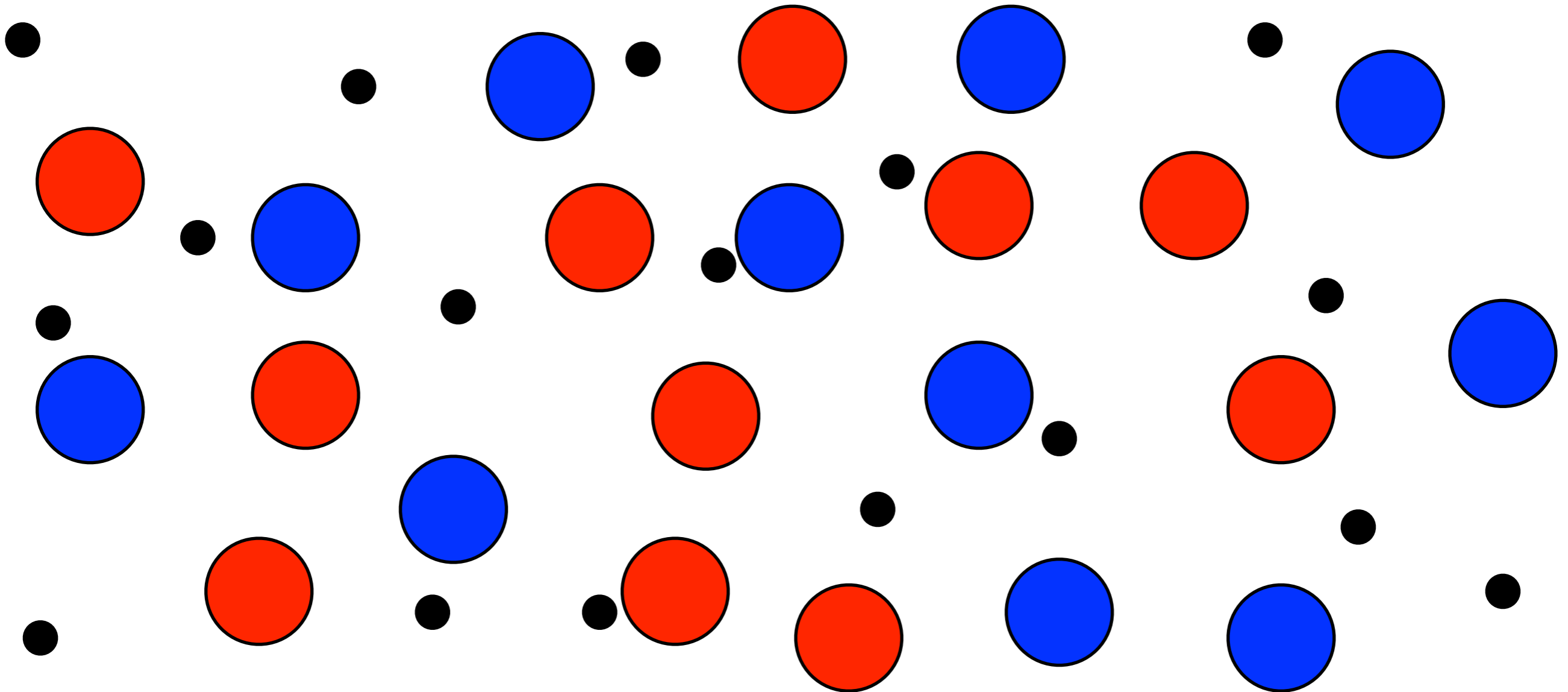
The first atoms

- A fraction ($1/10,000,000,000,000,000,000,000,000,000,000,000,000,000,000$) of a second after the BIG BANG electrons are created



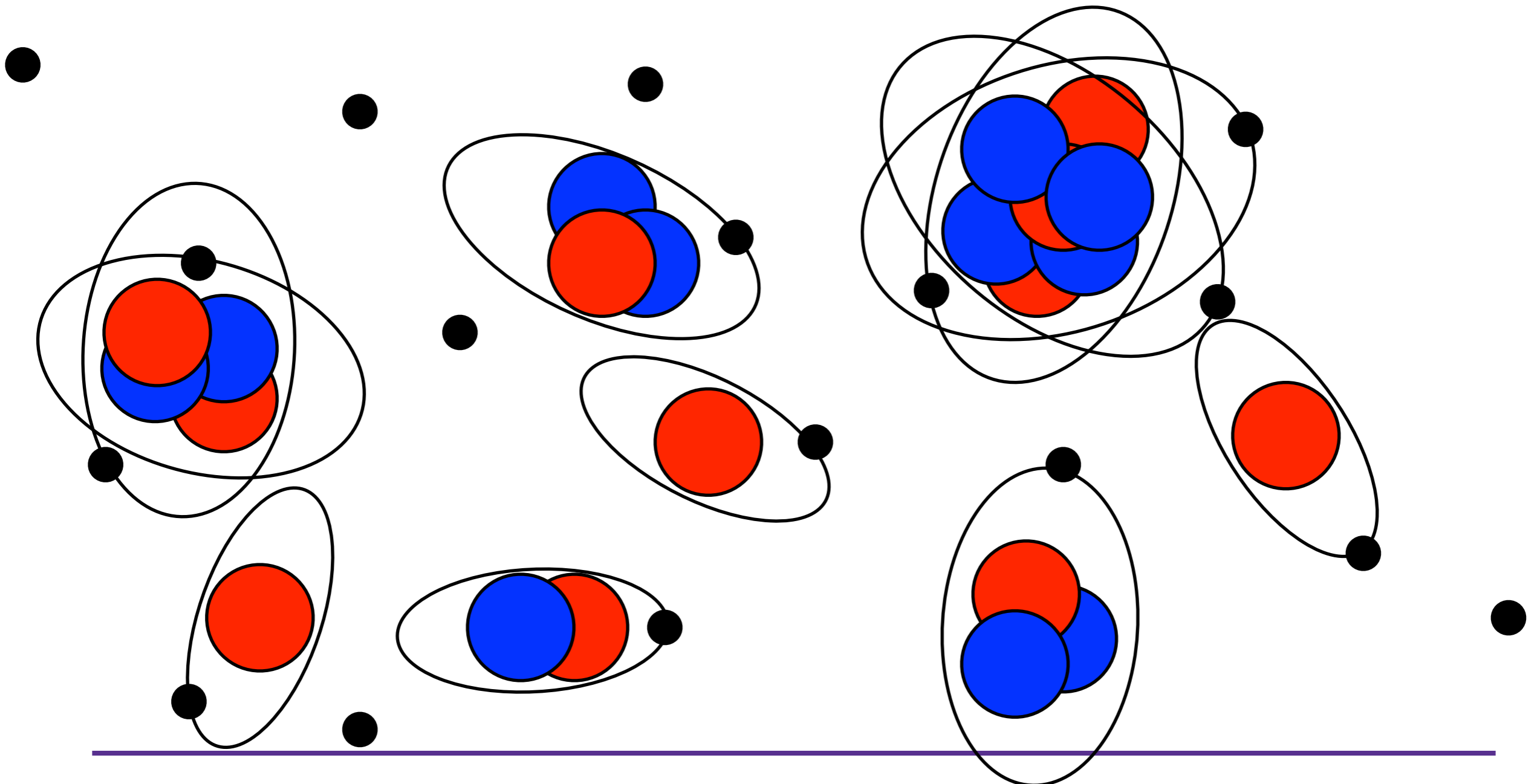
The first atoms

- One millionth of a second ($.000001$ seconds) after the BIG BANG protons and neutrons are formed



The first atoms

- 3 minutes after the BIG BANG the first atoms form



After the Big Bang

Periodic Table of the Elements © www.elementsdatabase.com

1 H	2 He											3 Li	4 Be	5 B	6 C	7 N	8 O	9 F	10 Ne												
11 Na	12 Mg											13 Al	14 Si	15 P	16 S	17 Cl	18 Ar														
19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr														
37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe														
55 Cs	56 Ba	57 La	72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn														
87 Fr	88 Ra	89 Ac	104 Unq	105 Unp	106 Unh	107 Uns	108 Uno	109 Une	110 Unn																						
																		58 Ce	59 Pr	60 Nd	61 Pm	62 Sm	63 Eu	64 Gd	65 Tb	66 Dy	67 Ho	68 Er	69 Tm	70 Yb	71 Lu
																		90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr

hydrogen
 alkali metals
 alkali earth metals
 transition metals
 poor metals
 nonmetals
 noble gases
 rare earth metals

What about everything else?



Calcium (Ca)



Sodium (Na)



Oxygen (O)
in water (H_2O)



Aluminum (Al)



Gold (Au)

Where does the rest of the Periodic Table come from?

Periodic Table of the Elements © www.elementsdatabase.com

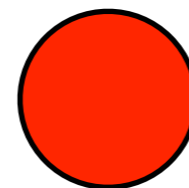
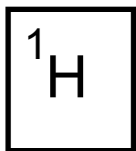
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Legend:

- hydrogen
- alkali metals
- alkali earth metals
- transition metals
- poor metals
- nonmetals
- noble gases
- rare earth metals

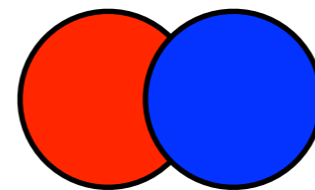
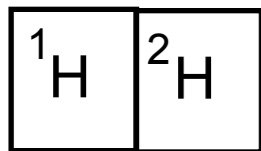
Beyond the Periodic Table . . .

- Adding or subtracting neutrons makes different *isotopes*
- A hydrogen nucleus is one **proton**
 - Adding one **neutron** to hydrogen makes the *isotope* deuterium
 - Then adding one **proton** makes the *isotope* ^3He
 - Then adding one **neutron** makes the *isotope* ^4He
- Total number of protons (“atomic number”) = Z
- Total number of neutrons = N
- $N + Z = A$, “atomic mass number” or number of nucleons



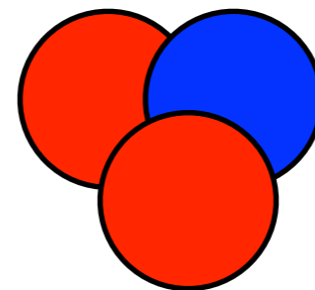
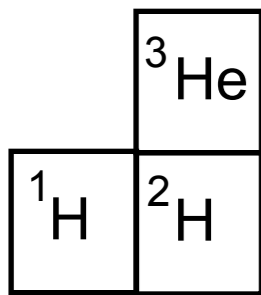
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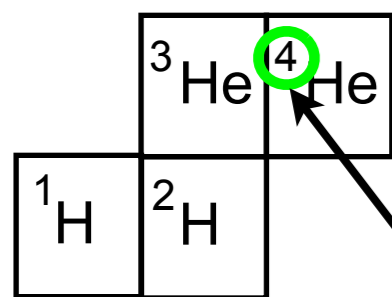
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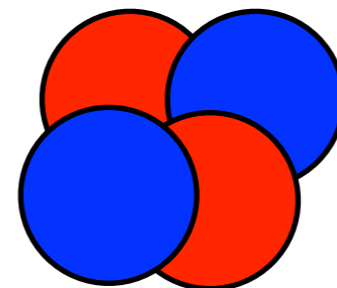


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protons + neutrons



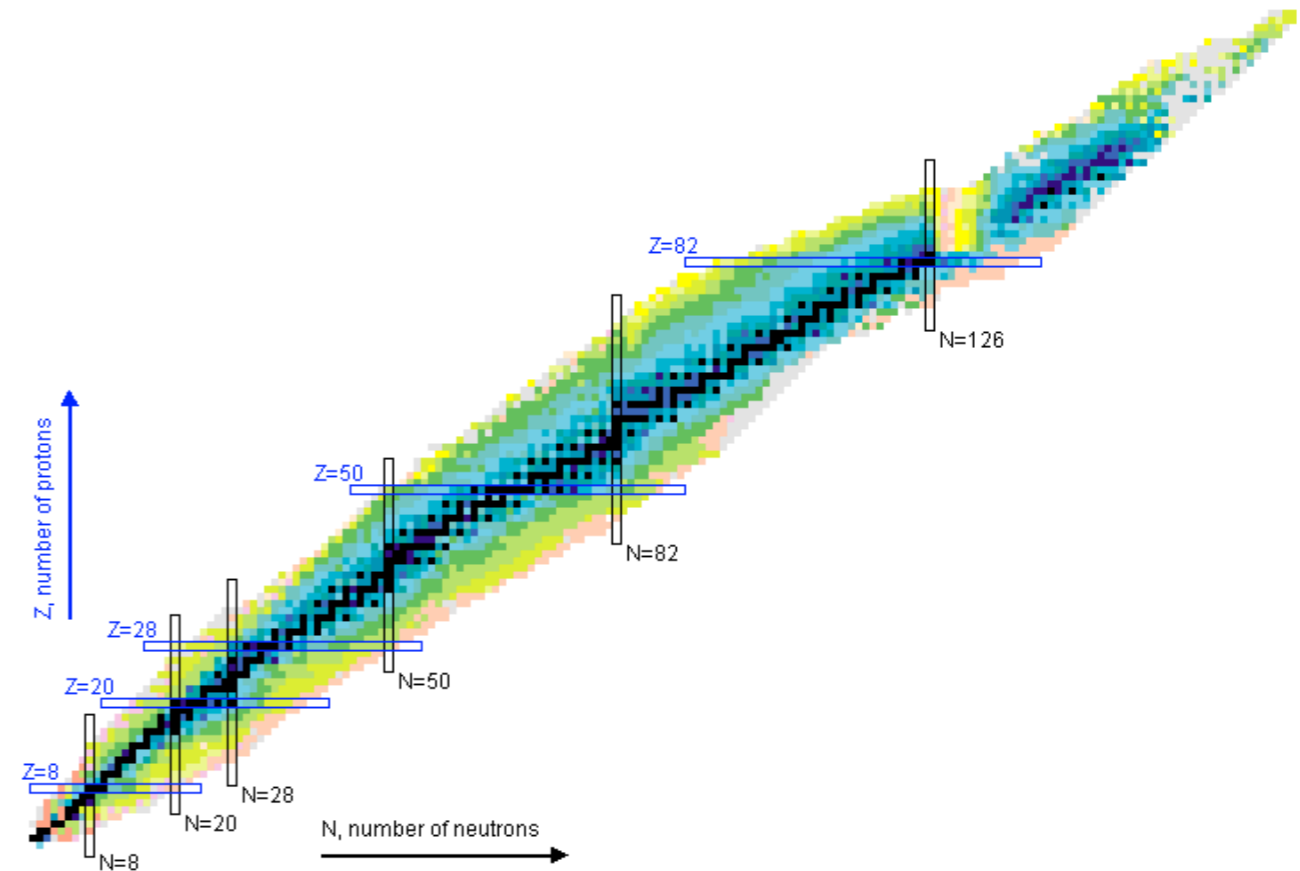
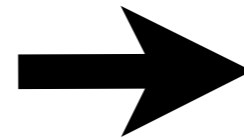
Beyond the Periodic Table . . .

- Keep adding protons and neutrons to make **thousands of isotopes**

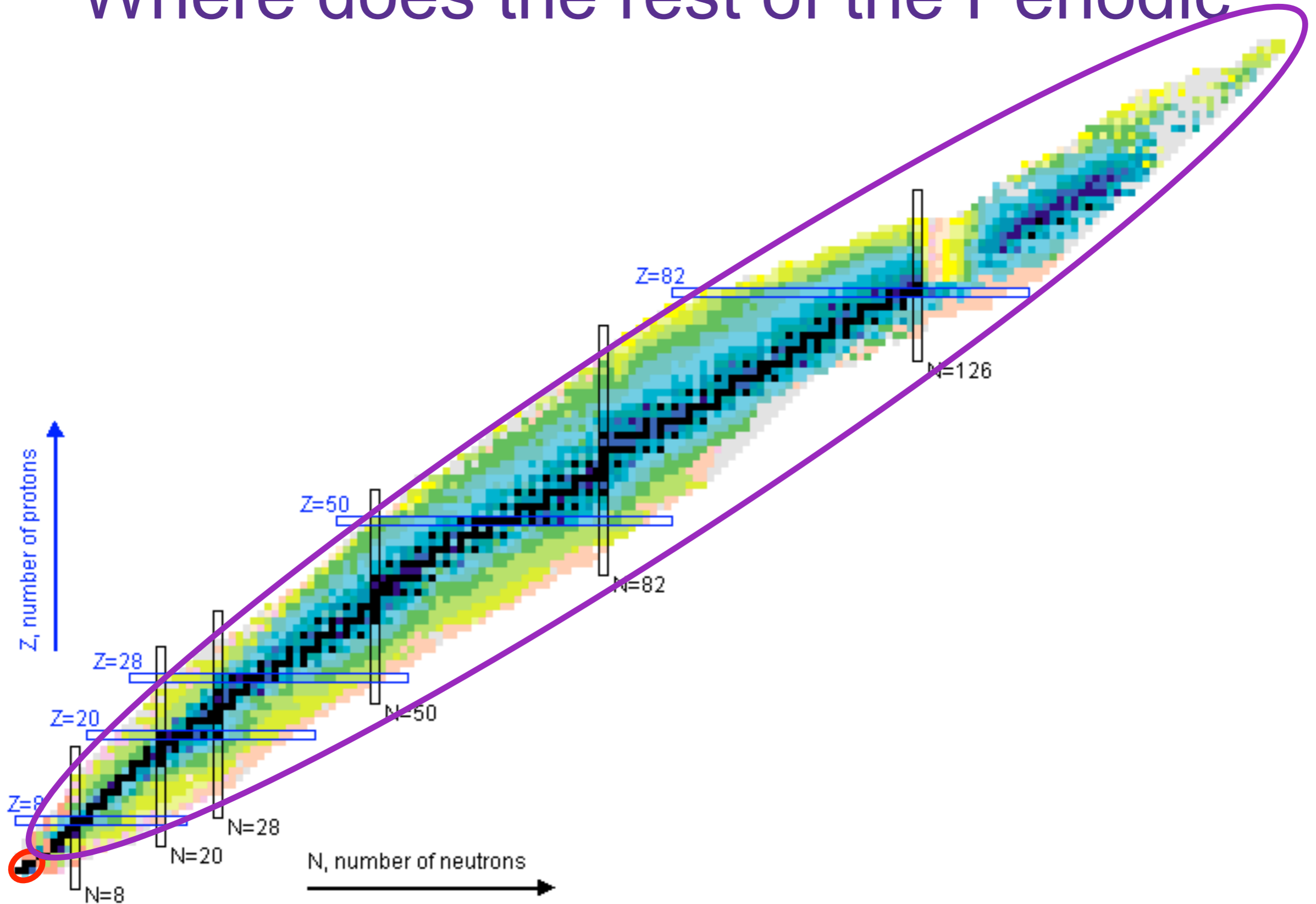
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■ alkali metals ■ nonmetals
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■ transition metals ■ rare earth metals



Where does the rest of the Periodic



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REVIEWS OF MODERN PHYSICS

VOLUME 29, NUMBER 4

OCTOBER, 1957

Synthesis of the Elements in Stars*

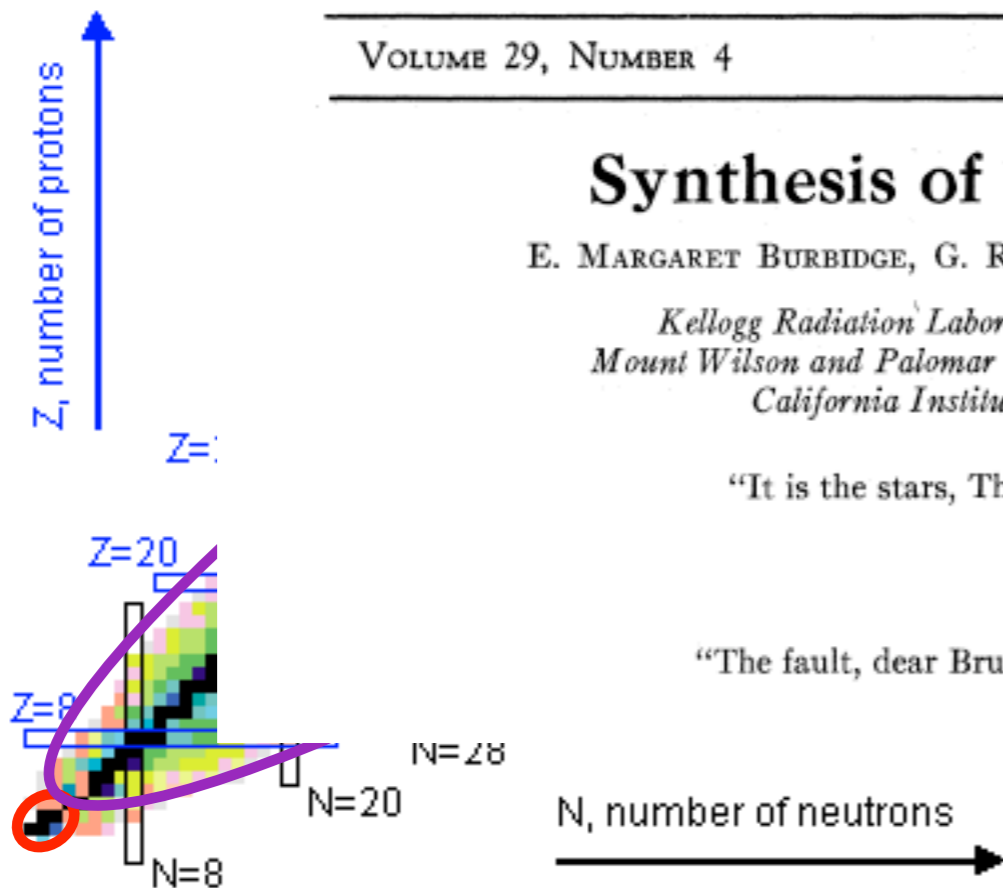
E. MARGARET BURBIDGE, G. R. BURBIDGE, WILLIAM A. FOWLER, AND F. HOYLE

*Kellogg Radiation Laboratory, California Institute of Technology, and
Mount Wilson and Palomar Observatories, Carnegie Institution of Washington,
California Institute of Technology, Pasadena, California*

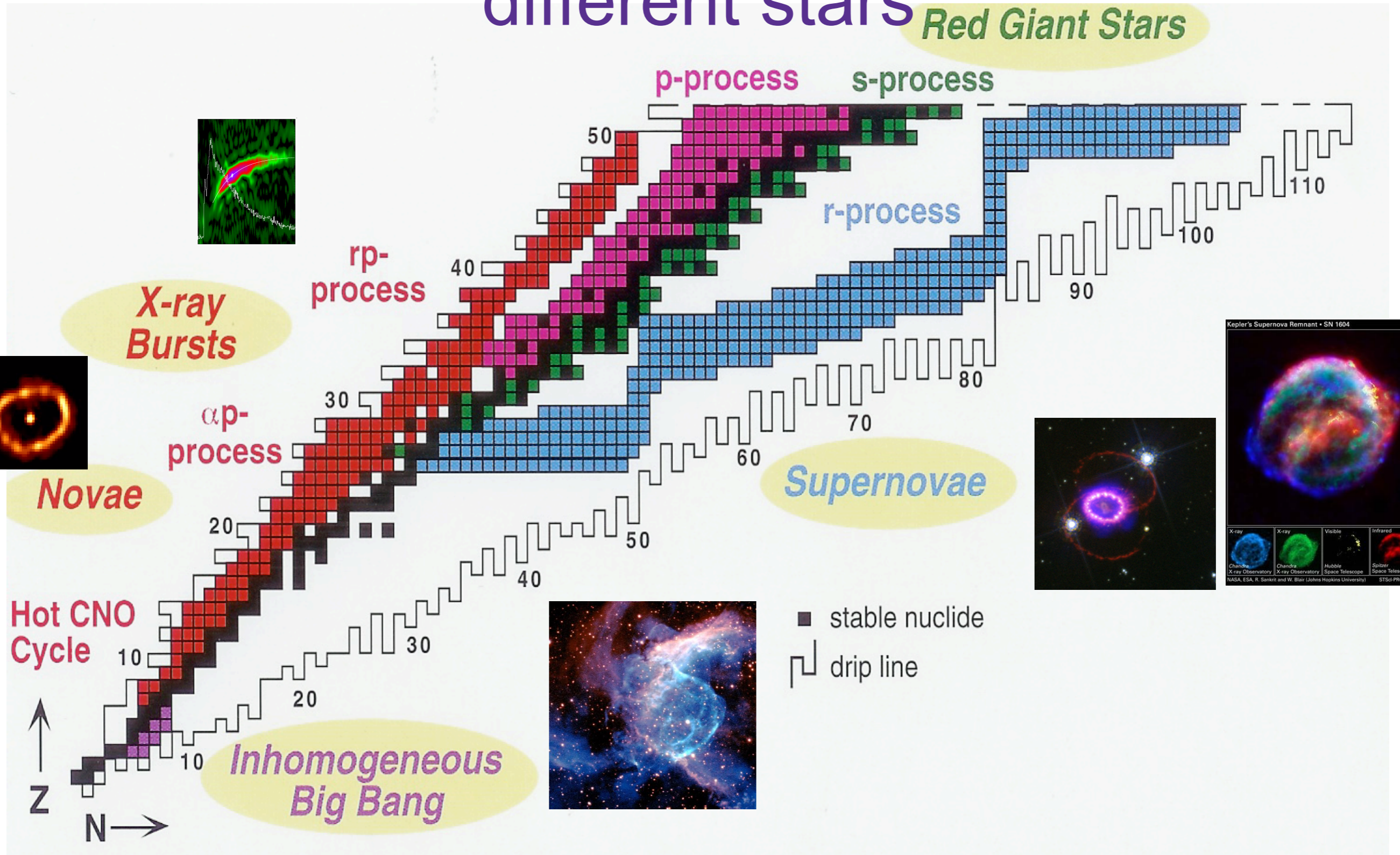
“It is the stars, The stars above us, govern our conditions”;
(King Lear, Act IV, Scene 3)

but perhaps

“The fault, dear Brutus, is not in our stars, But in ourselves,”
(Julius Caesar, Act I, Scene 2)



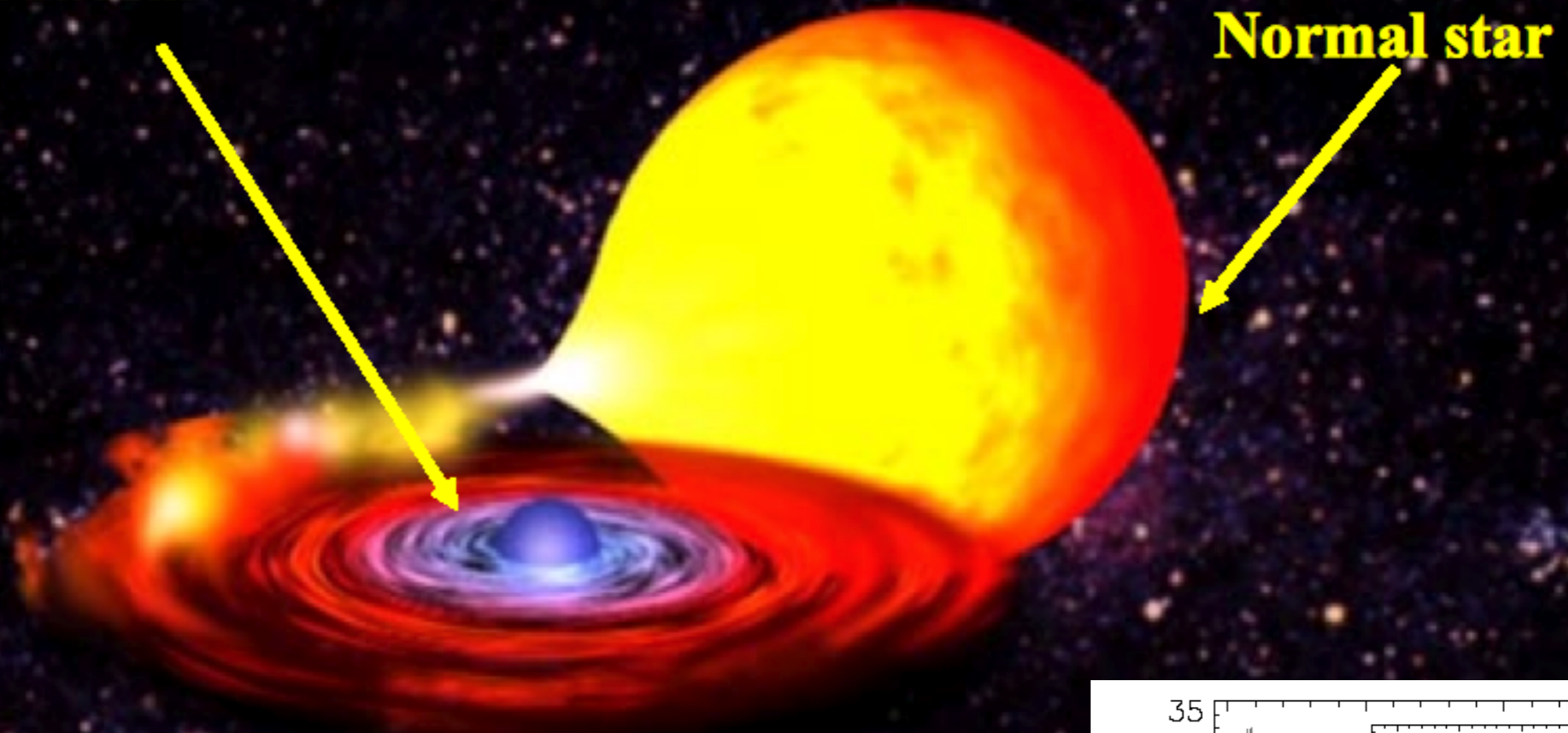
Different Nuclei are produced in different stars



Type I X-Ray Bursts (XRBs)

Neutron stars:

1.4 M_{\odot} , 10 km radius



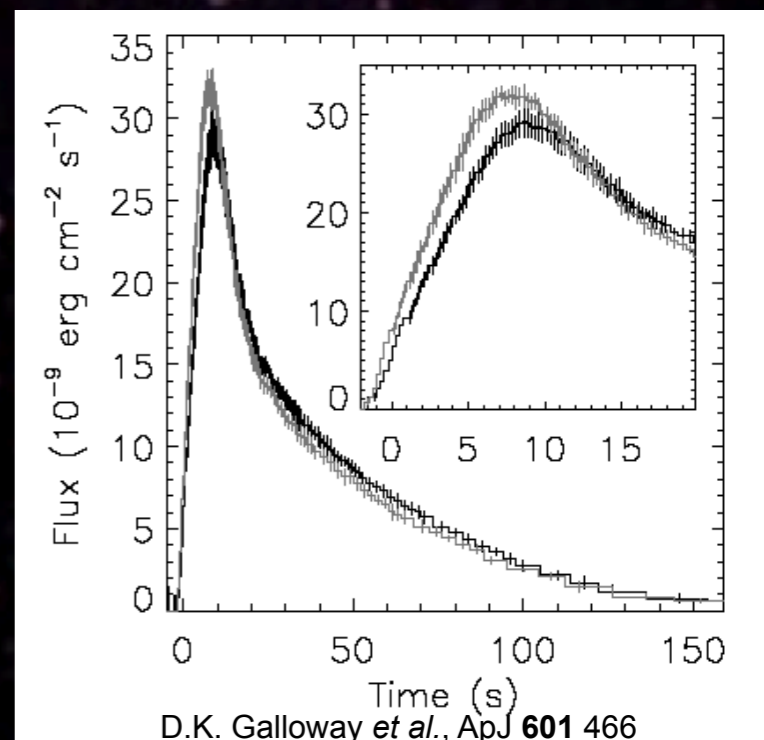
Accretion rate $\sim 10^{-8}/10^{-10} M_{\odot}/\text{year}$

Peak x-ray burst temperature $\sim 1.5 \text{ GK}$

Recurrence rate \sim hours to days

Burst duration of 10 – 100 s

Observed x-ray outburst $\sim 10^{39} - 10^{40}$ ergs



Neutron Stars

- Neutron stars are extremely compact, dense objects ($\rho \sim 10^{14}$ g/cm²)



Neutron Stars

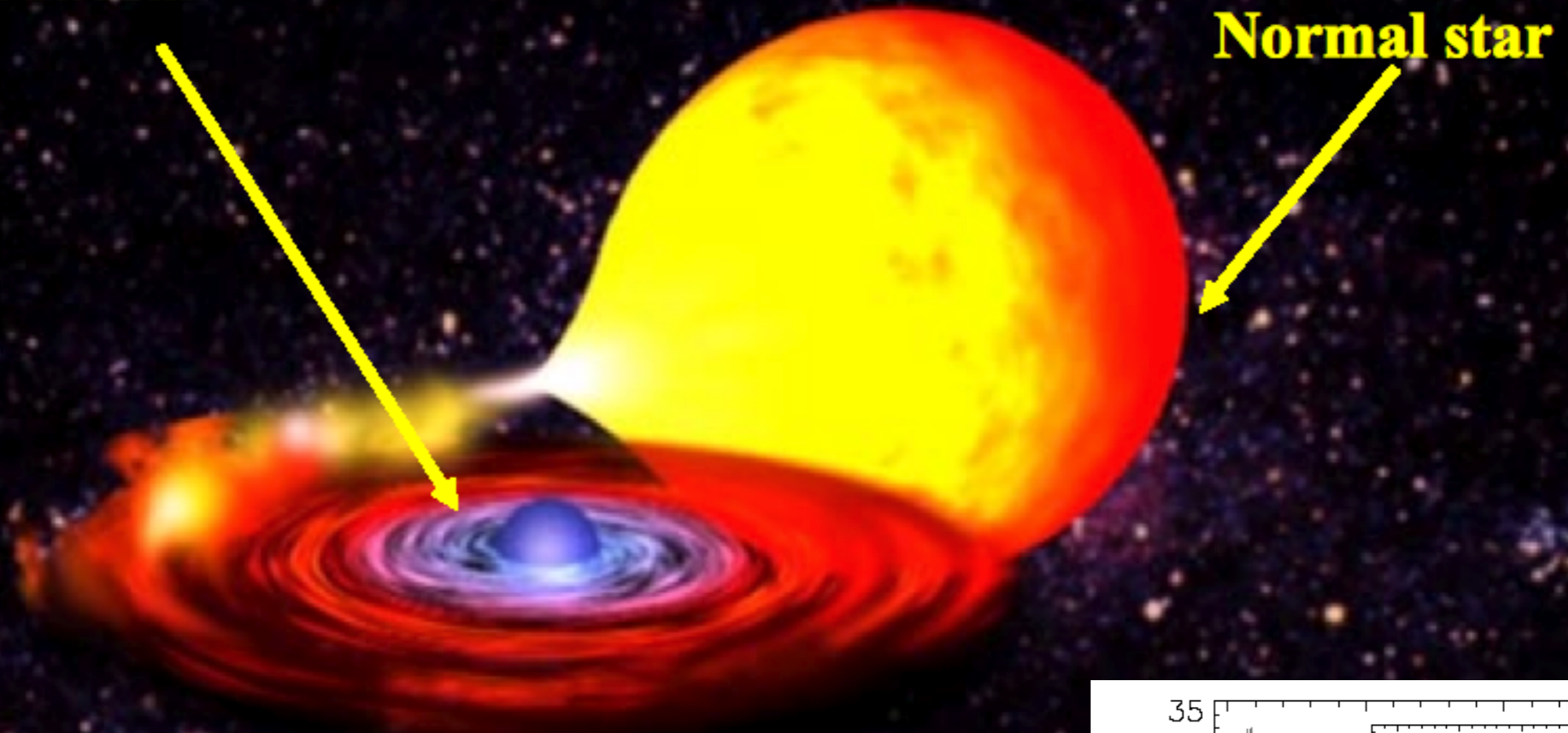
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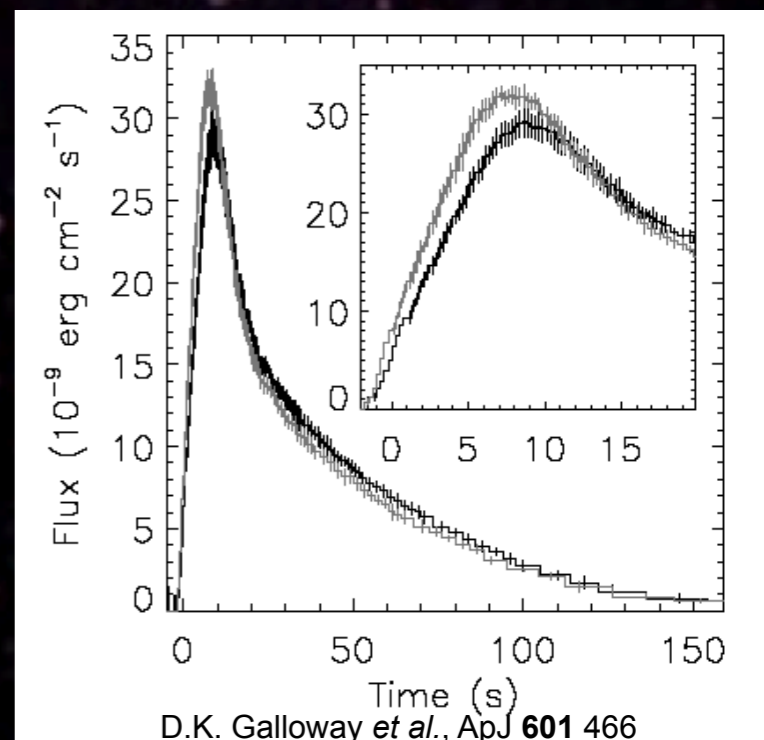
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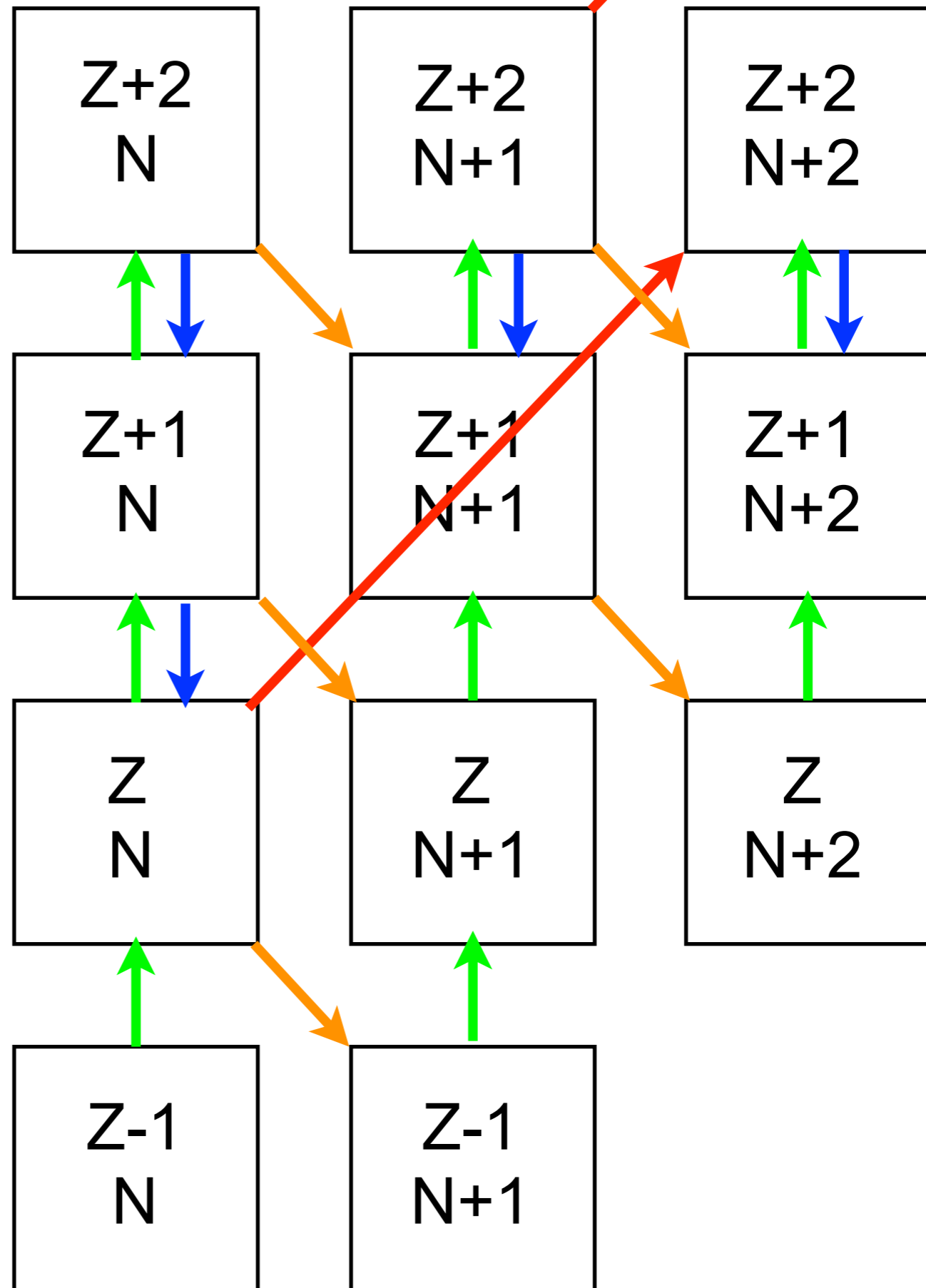
X-Ray Burst Nucleosynthesis

Proton capture:
(p, γ)

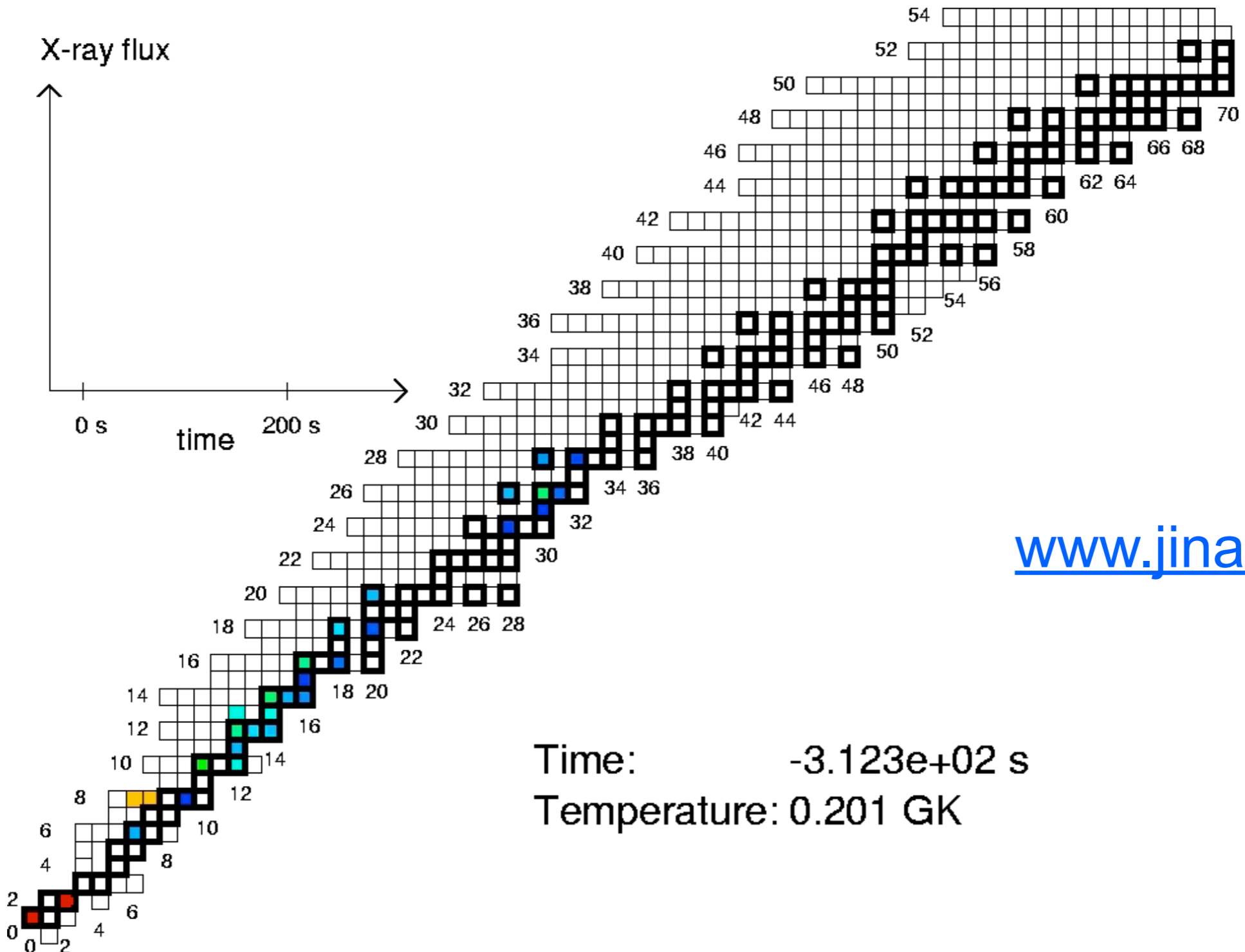
Proton decay:
(γ , p)

Beta decay:
(β^+)

Alpha capture:
(α , γ)
(α , p)



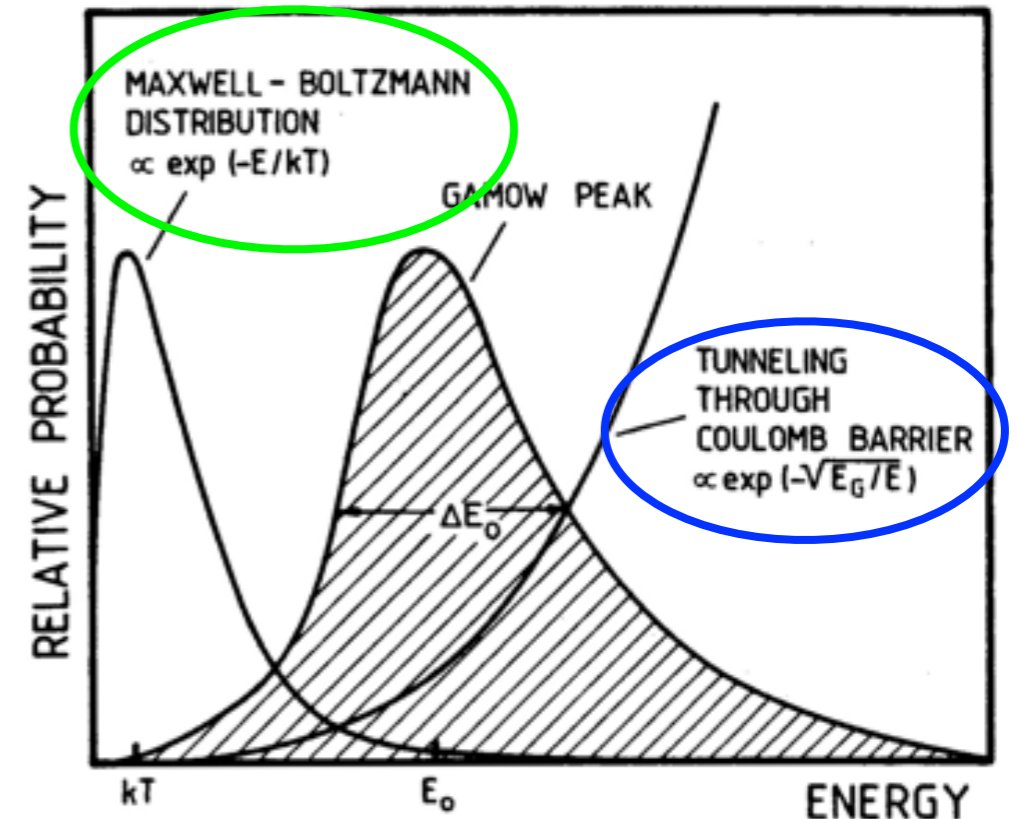
X-Ray Burst Nucleosynthesis



www.jinaweb.org

How do nuclei react?

- What is a **reaction rate**?
(i.e. What is the **probability** of two nuclei reacting in the stellar plasma?)
- Thermal distribution of nuclei in stellar plasma:
Maxwell-Boltzmann distribution
- The probability of the interaction between two nuclei:
nuclear cross section
- **Temperature dependent** – different temperatures in stars probe different energies in nucleus



Reaction Rates

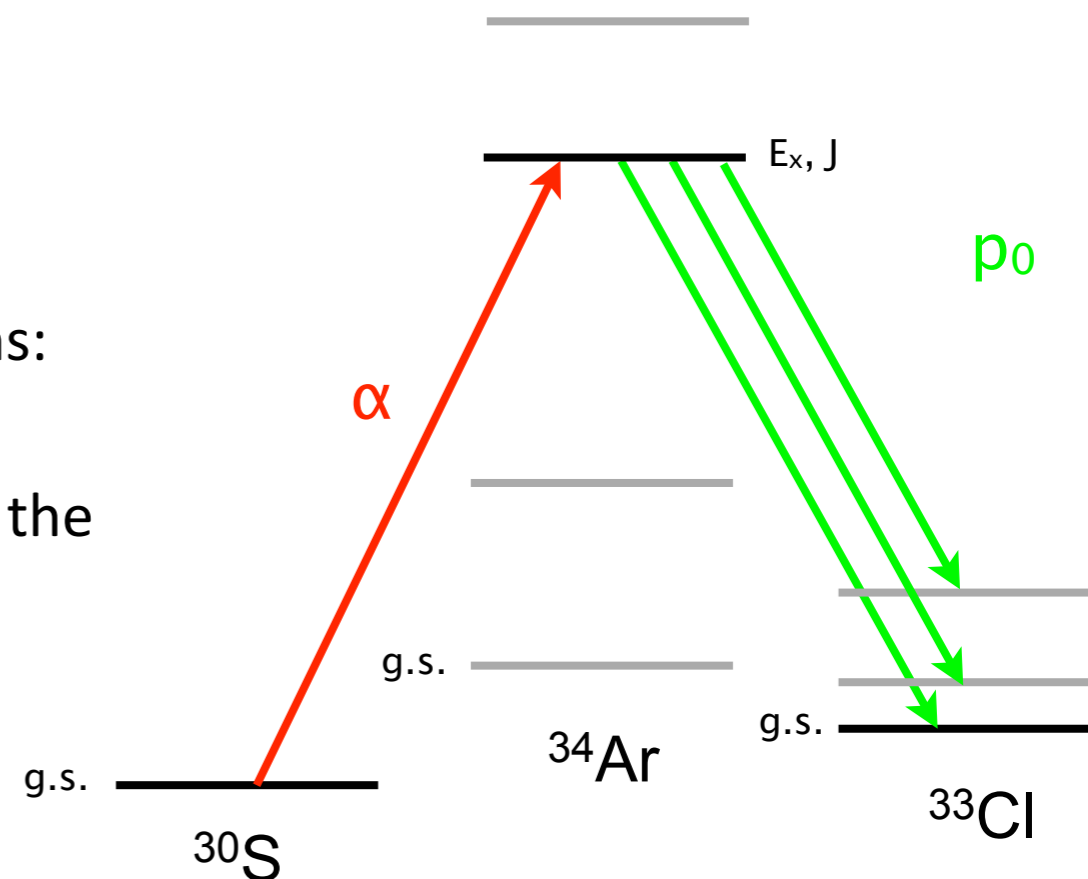
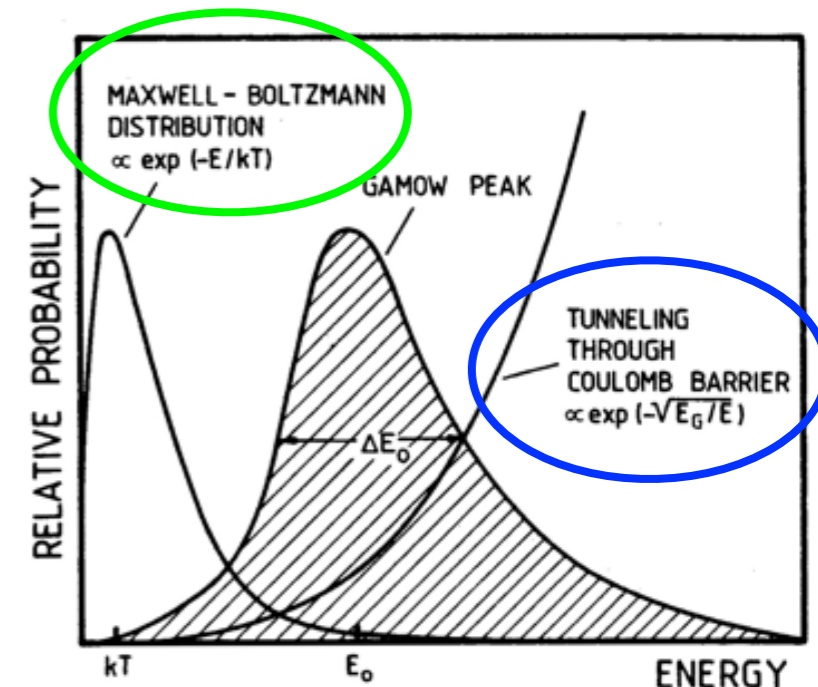
- Folding the **Maxwell-Boltzmann distribution** in with the **nuclear cross section** gives the **reaction rate**

- For resonant reaction rates:

- $\propto \exp(-E)$
- \propto nuclear spin, J
- \propto nuclear widths, Γ

- Two ways to study reactions rates and cross sections:
 - Directly- measuring the reaction rate itself
 - Indirectly- determining different components of the reaction rate

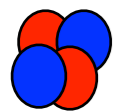
- e.g. $^{30}\text{S} (\alpha, p)^{33}\text{Cl}$



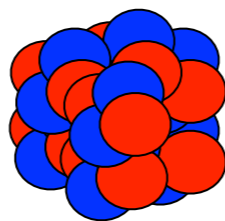
Studying Nuclear Reactions in the Laboratory

- Using accelerators with different types of detectors we can measure what nuclear reactions happen in stars
- A particle beam is accelerated and impinges on a target
- Outgoing particles are detected

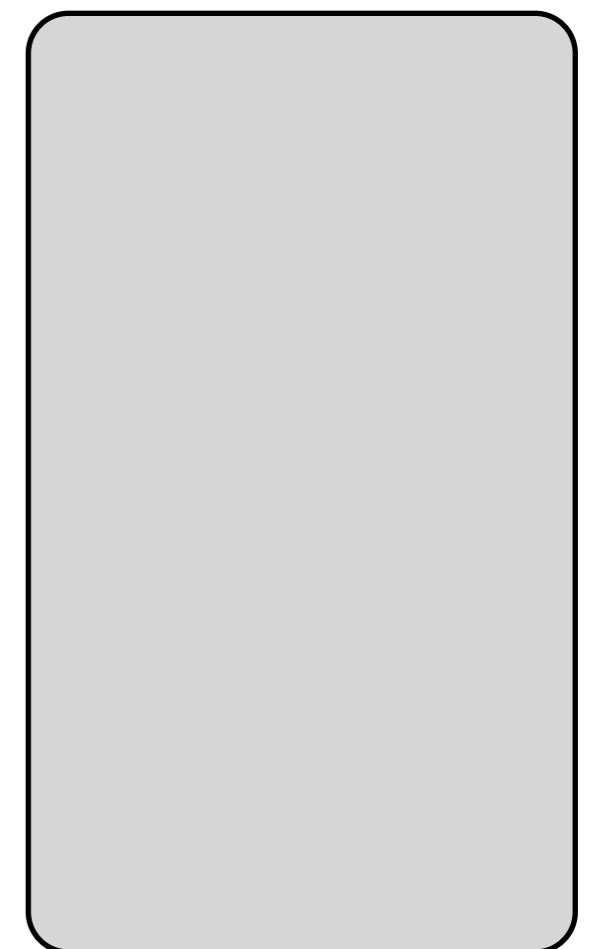
DETECTOR



beam →

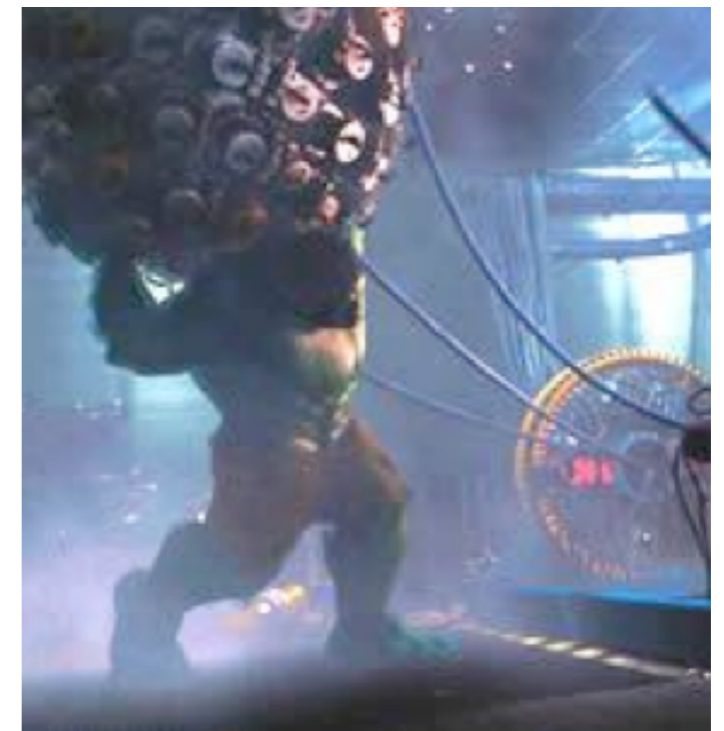
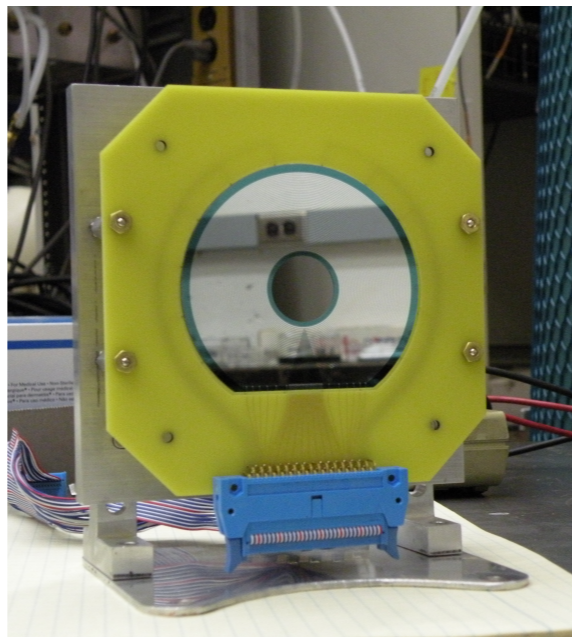


target



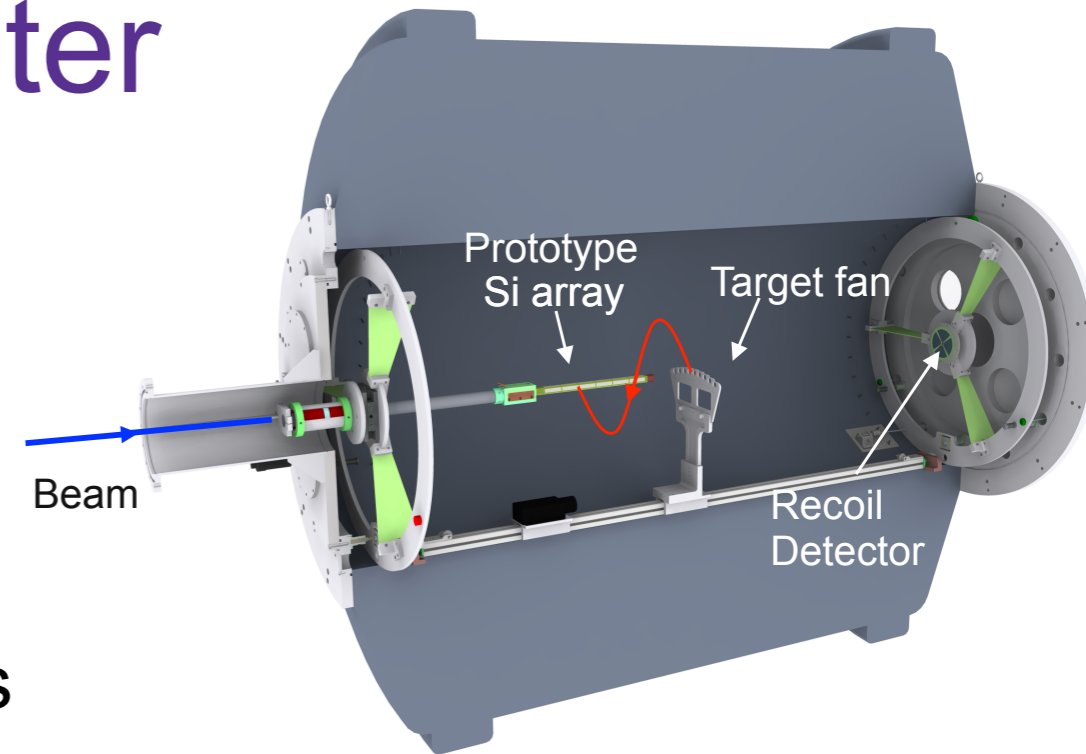
Studying Nuclear Reactions in the Laboratory

- Charged particles can be manipulated by magnetic fields and separated by
 - charge
 - mass
 - energy
- Detected using
 - ionization chambers
 - silicon detectors
 - CsI detectors
 - gamma detectors

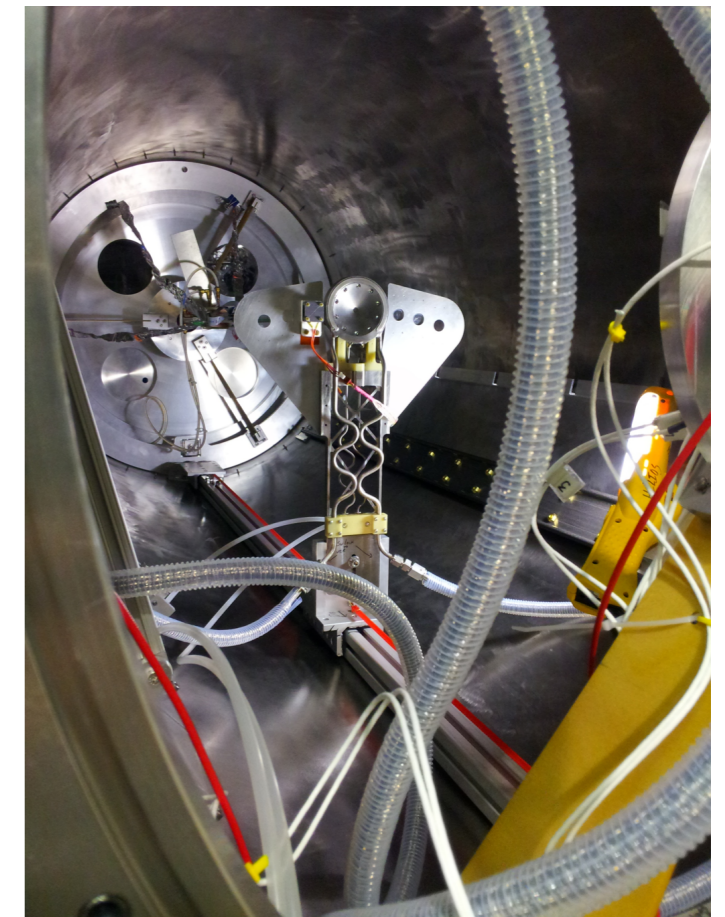
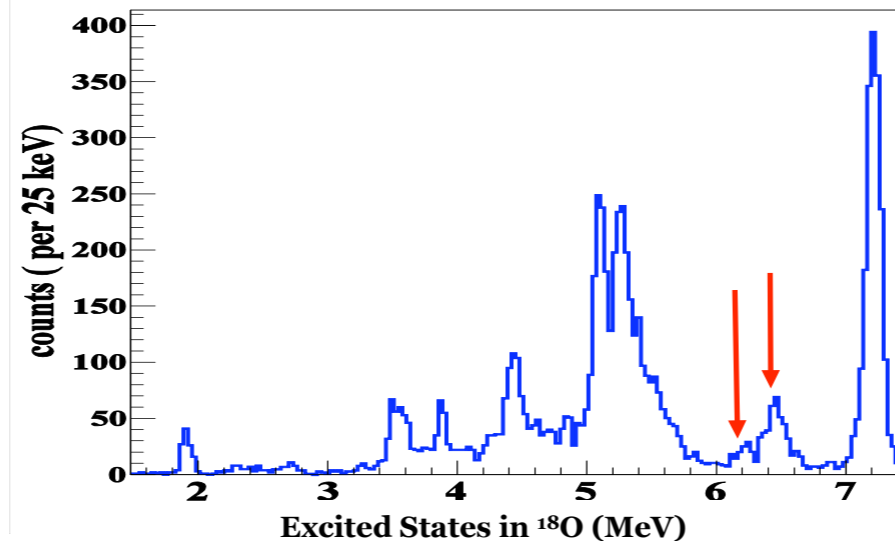


HELICAL Orbit Spectrometer

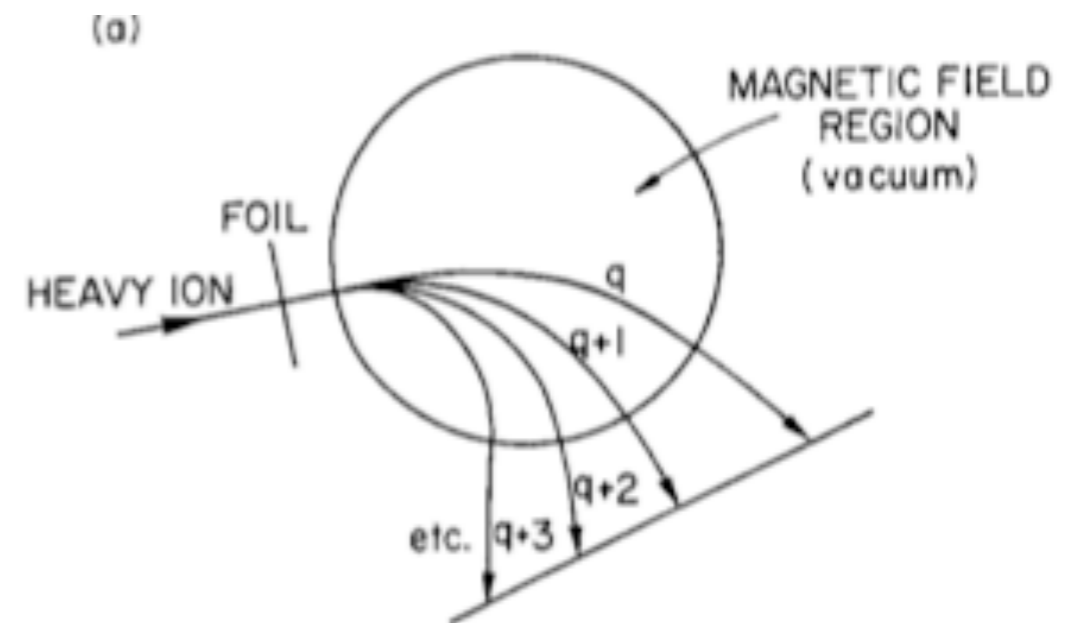
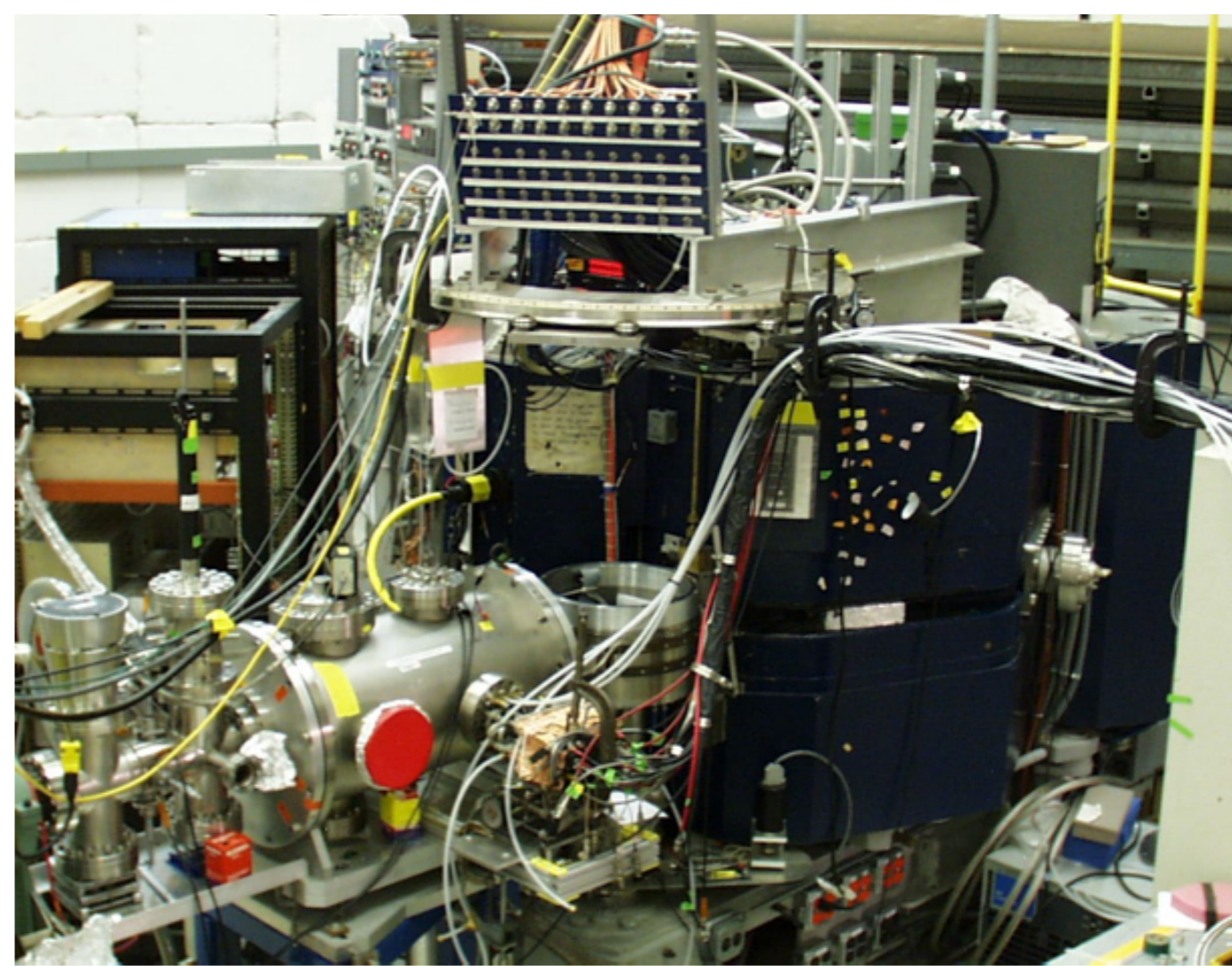
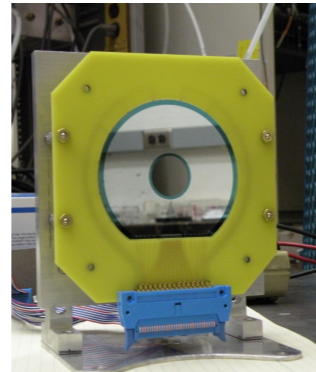
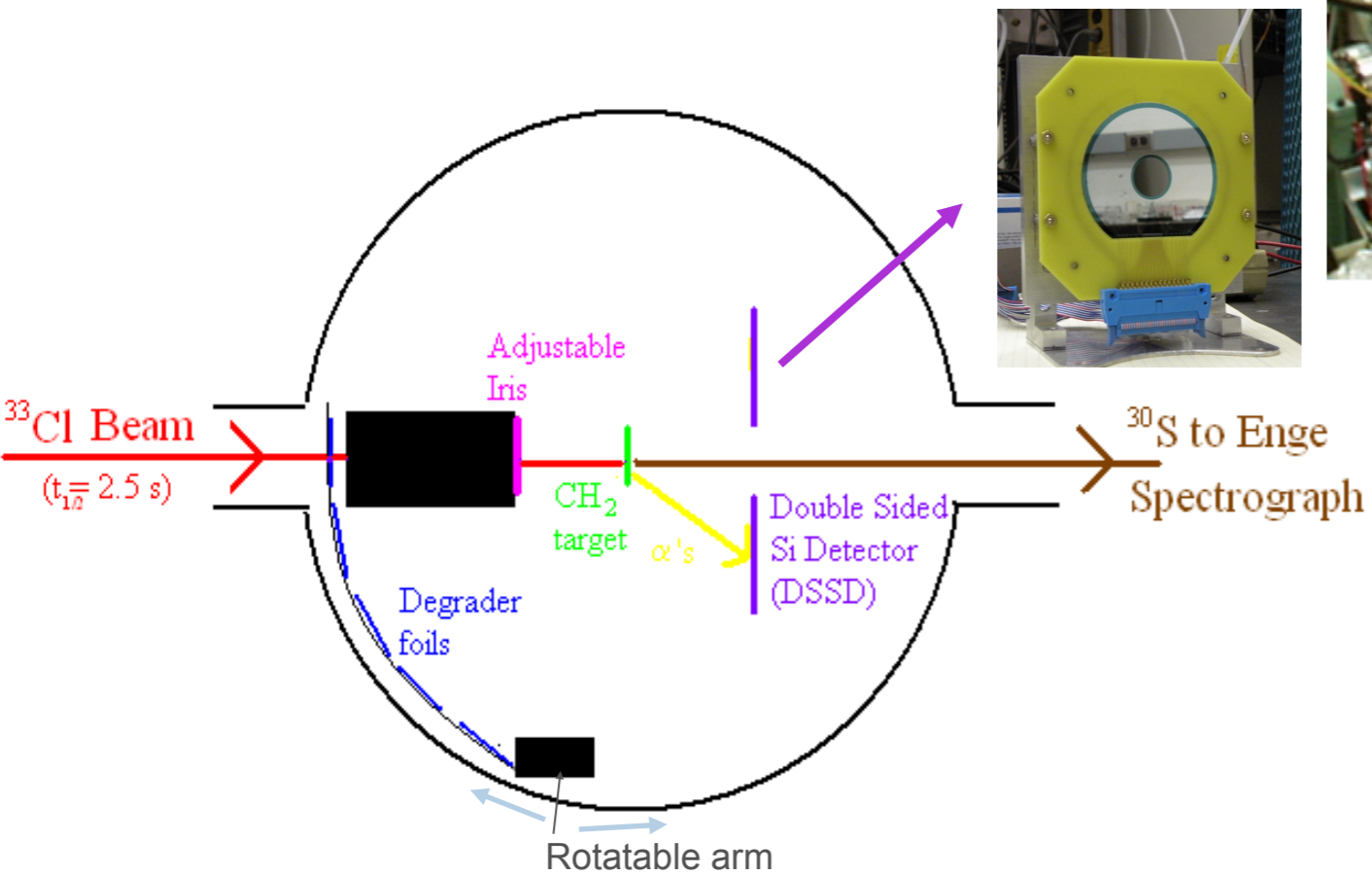
- Beam of radioactive nuclei directed through center of solenoid
- Impinges on a target of light nuclei (e.g. hydrogen, helium, etc.)
- Reaction products measured by detectors
- Reaction products tell us:
 - excitation energy levels
 - spins of levels
 - reaction rate information



States in ^{18}O from
 $^{14}\text{O}(^6\text{Li},d)^{18}\text{O}$



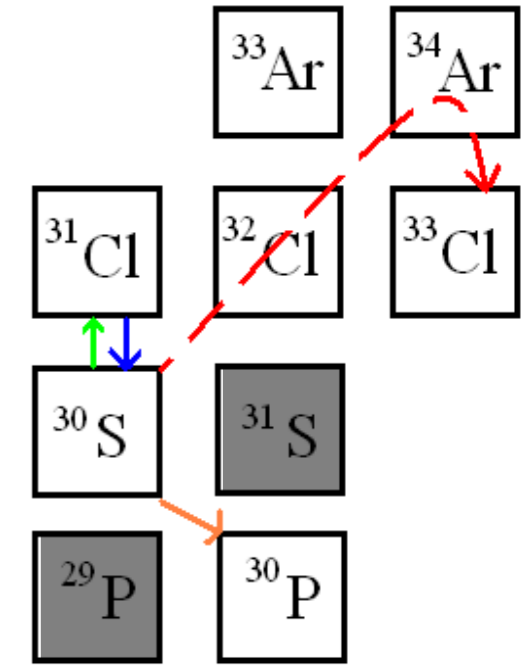
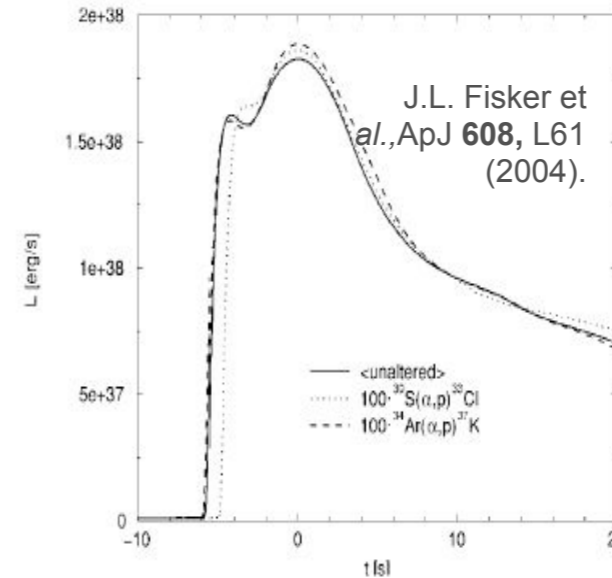
$^{33}\text{Cl}(p,\alpha)^{30}\text{S}$ Measurement



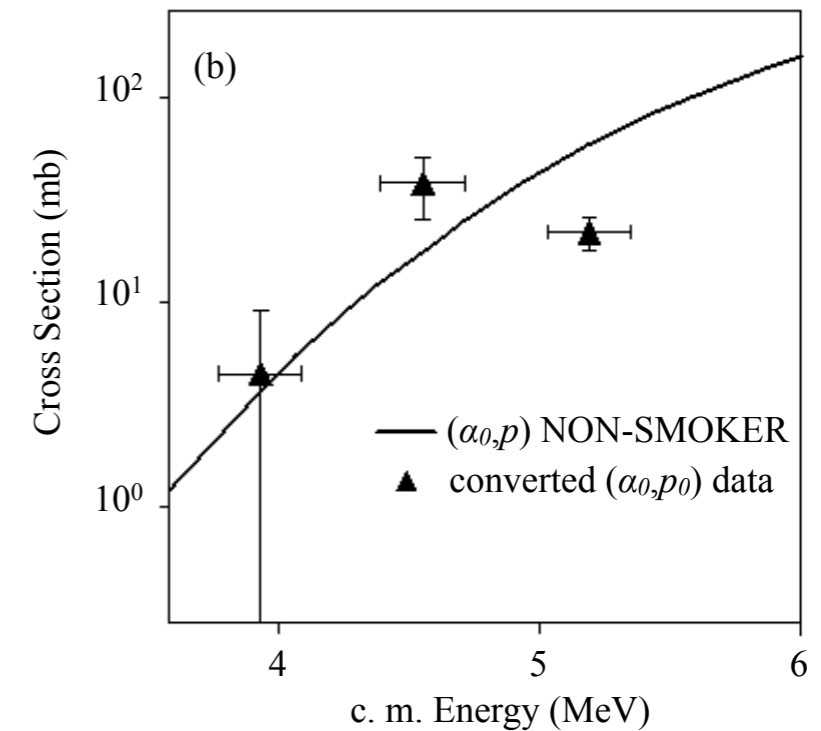
(α, p) -process waiting points:

$^{30}\text{S}(\alpha, p)^{33}\text{Cl}$ Measurement

- (α, p) reactions on waiting points (^{22}Mg , ^{26}Si , ^{30}S , and ^{34}Ar) may have significant effects on type I X-ray bursts
 - final elemental abundances
 - energy generation
 - double-peaked luminosity profiles
- Measured cross sections (probability of reacting) larger than theoretical predictions:
 - reaction rate is bigger!



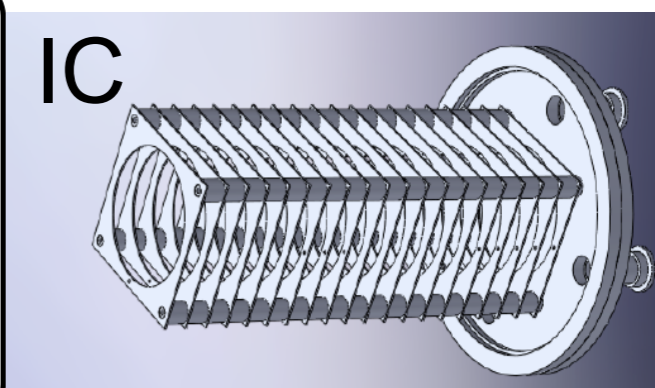
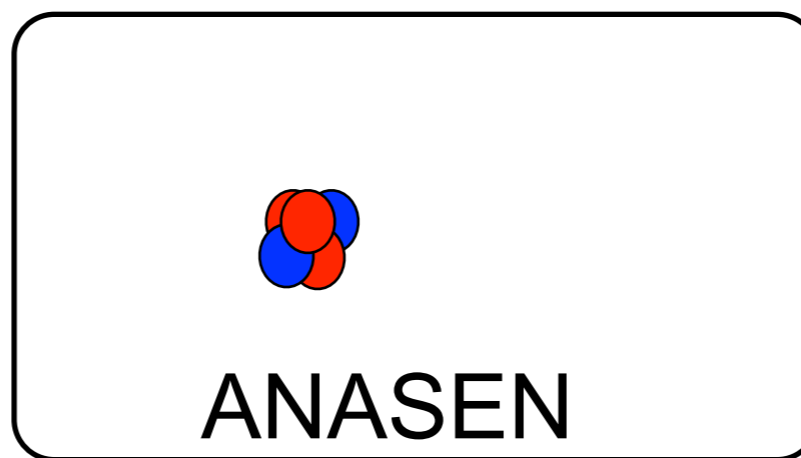
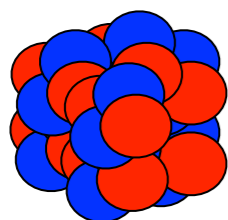
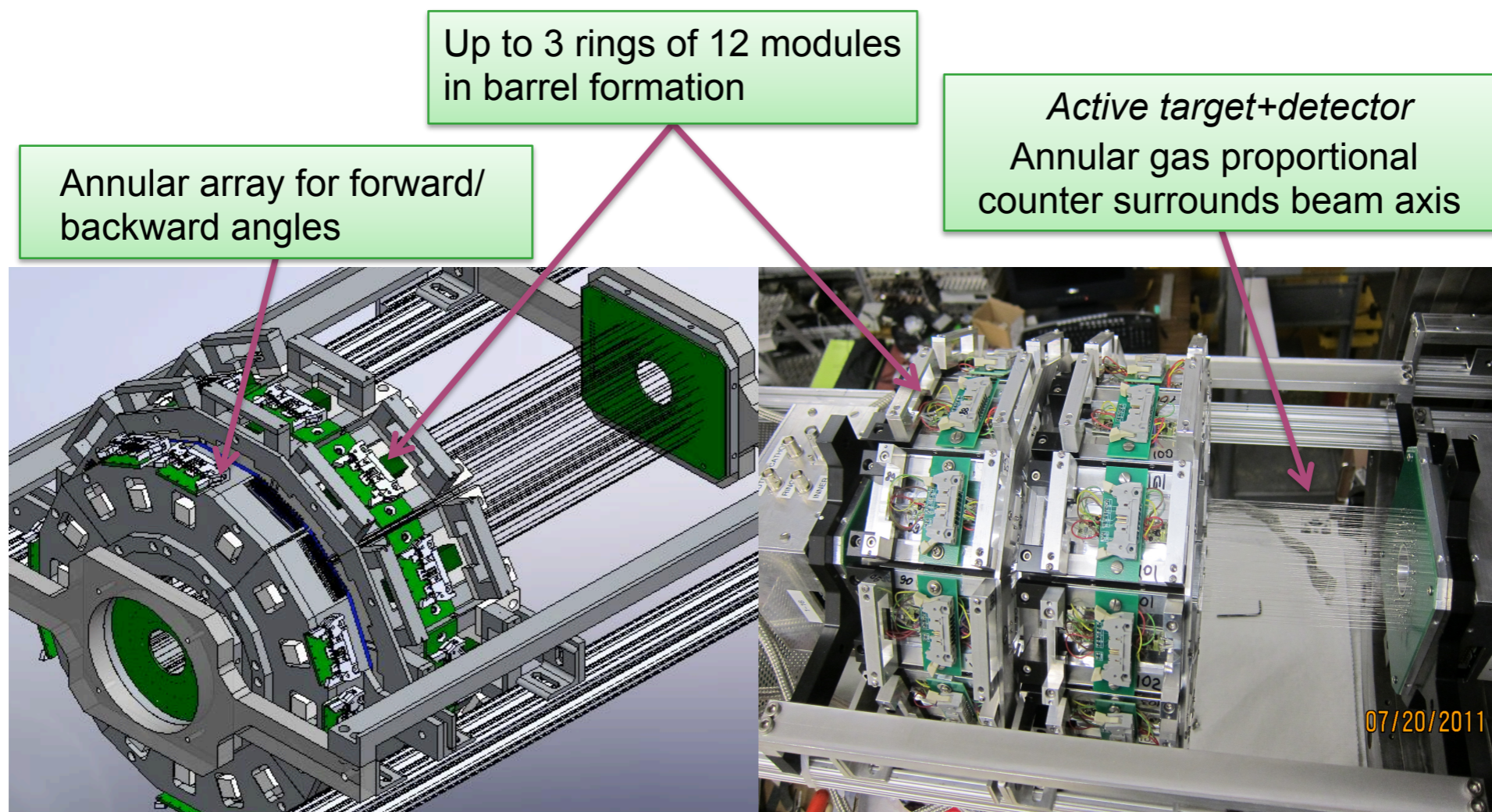
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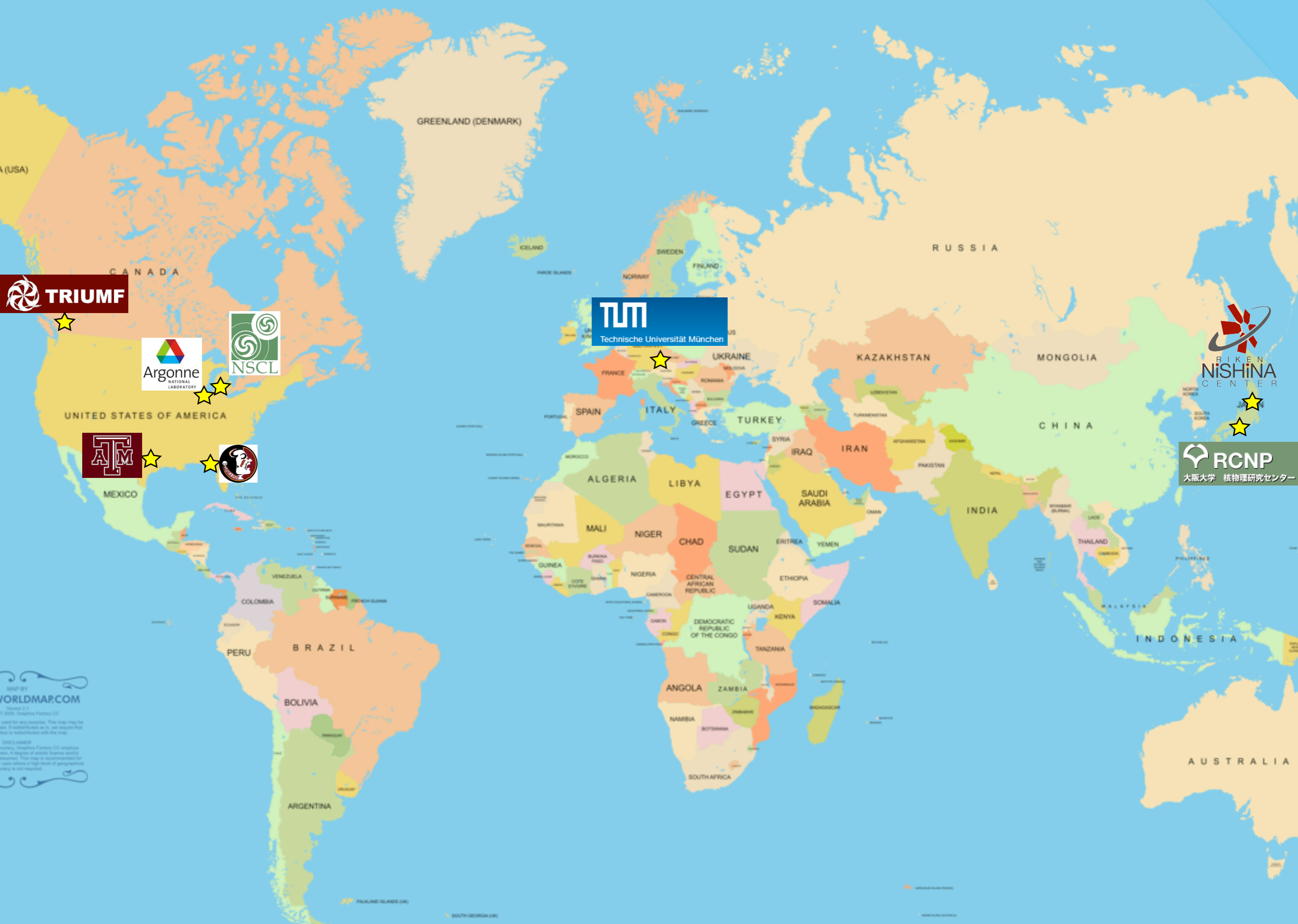


C.M. Deibel et al, submitted (2011).

Experiments for X-ray bursts

Array for Nuclear Astrophysics Studies with Exotic Nuclei (ANASEN)



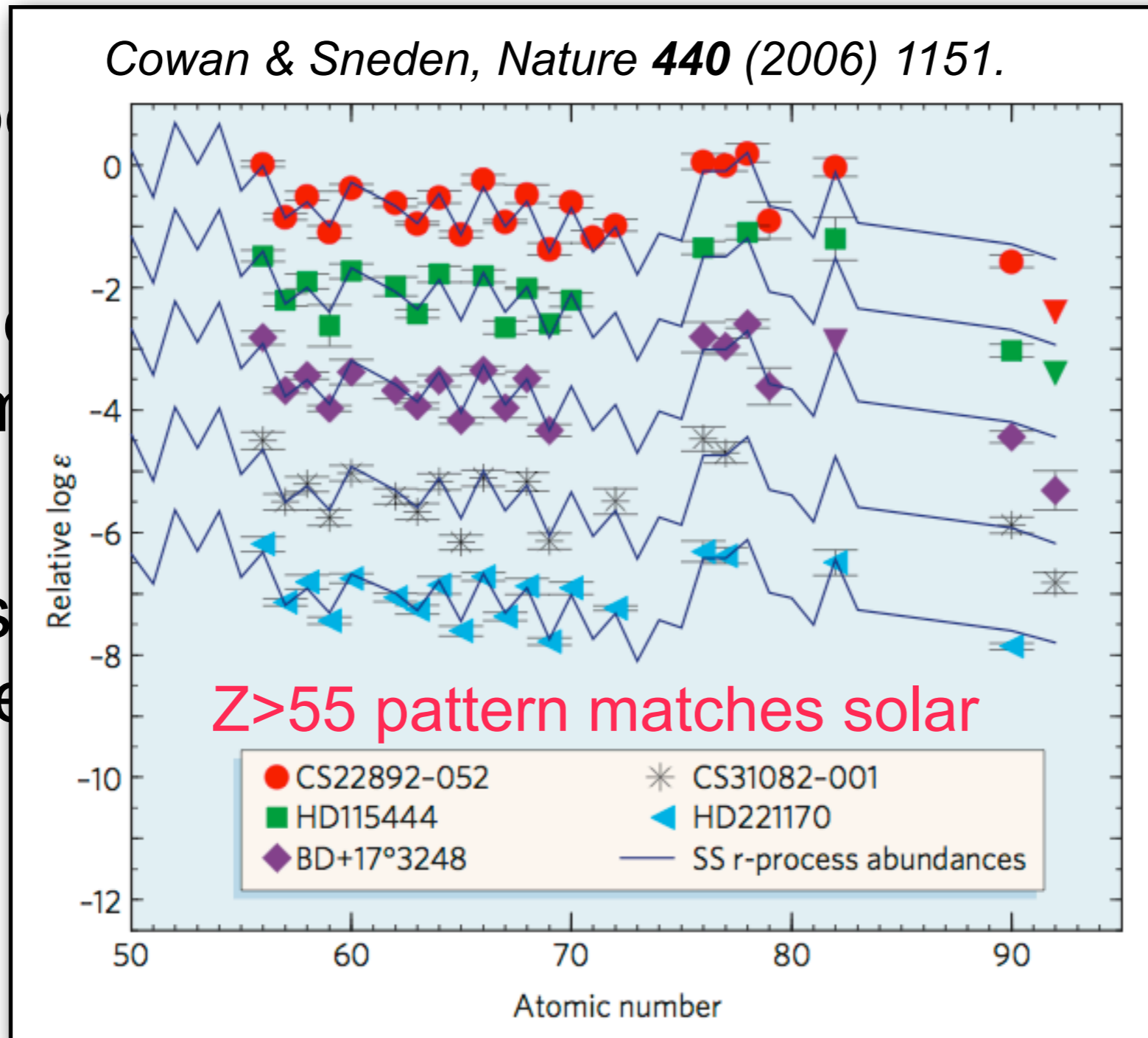


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Embarrassing truths about NA

a.k.a. why we stay employed

- Sup
- We
- elem
- Mos
- have



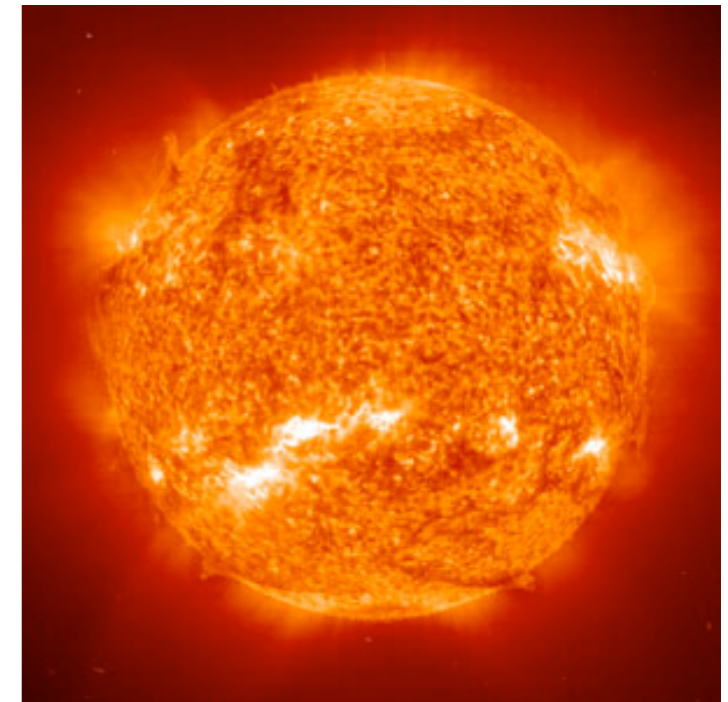
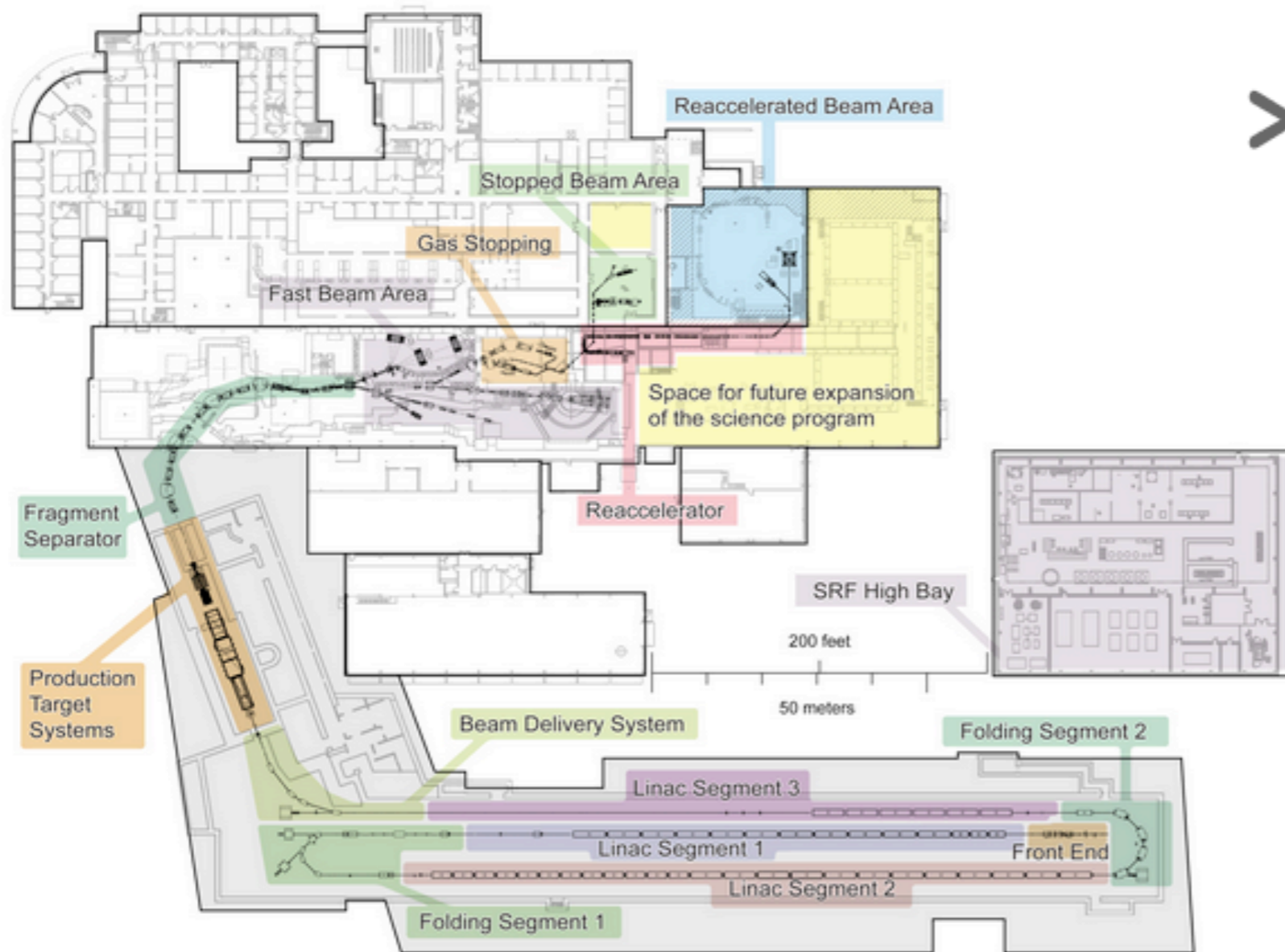
Embarrassing truths about NA

a.k.a. why we stay employed

- Supernova models don't explode
- We don't know where all the heavy elements are made
- Most reactions that happen in stars have not been studied
- But the future is bright . . .

Embarrassing truths about NA

a.k.a. why we stay employed



Layout of the accelerator and experimental systems and the experimental areas of the Facility for Rare Isotope Beams.

Image 2 of 3



Thanks!



Rascoe

Blackmon



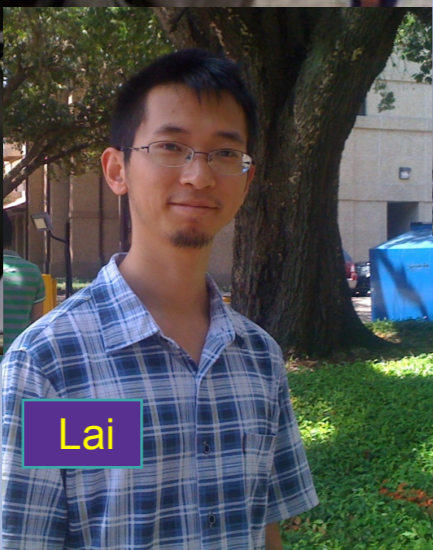
Williams



Gardiner



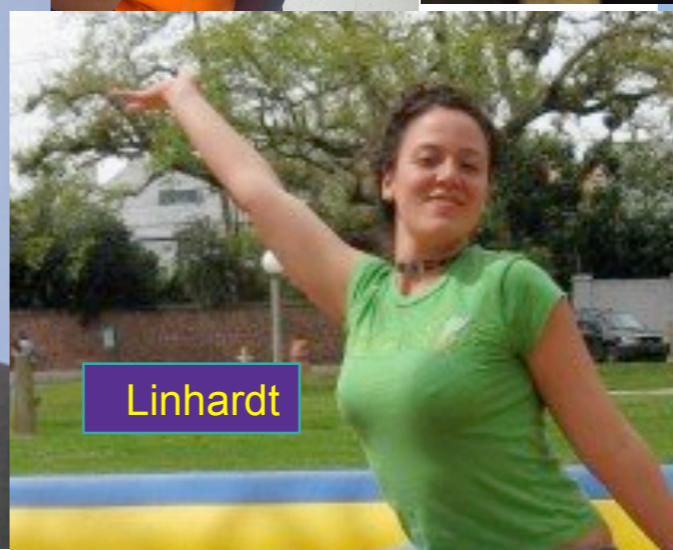
Zganjar



Lai



Macon



Linhardt



Afanasieva



Lauer